

Using a non-reducing sugar in the green synthesis of gold and silver nanoparticles by the chemical reduction method

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Abstract

The use of more biocompatible and renewable chemical compounds to obtain metal nanoparticles with desired properties and characteristics becomes an alternative route for the reduction of environmental risks and to minimize the incompatibility of these structures when interacting with biological models for their possible application in the health area. The purpose of this research was focused on the use of sucrose, as a reducing agent for gold and silver nanoparticles using different volumes of sodium hydroxide. The nanoparticles obtained were characterized by UV-visible spectrometry, transmission electron microscopy TEM and Fourier transform infrared spectroscopy FTIR, which allowed to determine the surface plasmon resonance, experimental and theoretical particle sizes, morphology and structural changes in the reducing agent, as well as the influence of sodium hydroxide in the synthesis process. The results obtained confirm the generation of gold and silver nanoparticles by the previous formation of reducing sugars and oxidation of the functional group from glucose to salts of carboxylic acid.

Keywords: metal nanoparticles; sucrose; hydrolysis; reduction oxide reaction; salt of carboxylic acid.

Uso de una azúcar no reductora en la síntesis verde de nanopartículas de oro y plata por medio de reducción química

Resumen

El uso de compuestos químicos más biocompatibles y renovables para la obtención de nanopartículas metálicas con propiedades y características deseadas, se convierte en una ruta alternativa para la reducción de riesgos ambientales y del grado de incompatibilidad de estas estructuras al interactuar con modelos biológicos para su posible aplicación en el área de la salud. El propósito de este trabajo se centró en el uso de sacarosa, como agente reductor de nanopartículas de oro y plata al emplear diferentes volúmenes de hidróxido de sodio. Las nanopartículas obtenidas fueron caracterizadas mediante espectrometría UV-visible, microscopía electrónica de transmisión TEM y espectroscopia infrarroja por transformada de Fourier FTIR, la cual permitió determinar los plasmones de resonancia superficial, tamaños de partícula experimentales y teóricos, morfología y cambios estructurales en el agente reductor, así como la influencia del hidróxido de sodio en el proceso de síntesis. Los resultados obtenidos confirman la formación de nanopartículas de oro y plata mediante la previa formación de azúcares reductores. Así mismo, la oxidación del grupo funcional de la glucosa a sales de ácido carboxílico.

Palabras clave: nanopartículas metálicas; sacarosa, hidrólisis; óxido reducción; sales de ácido carboxílico.

1. Introduction

Nanoparticles have been potentiated for the research and development of biosensors, catalytic processes, and biomedical applications, among others. The interest in these

materials is due to the fact that is possible to obtain new properties and characteristics by modifying their size and structure depending on the production process employed [1-3]. Currently, different physical and chemical processes are used to synthesize metal nanoparticles (MNPs) to produce

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particles with characteristics and properties related to the application area. Two kinds of methodologies are used to obtain nanostructures, i) top-down strategies, in which dimensionally bulk materials are collapsed little-by-little by manual or automated processes until reaching nanoscale [4], and ii) bottom-up where the molecular modification of a precursor agent in order to obtain atoms, which by nucleation processes result in monodisperse nanostructures [5]. However, most of these methods are often expensive, labor-intensive, potentially hazardous to the environment and highly cytotoxic [6]. Therefore, there is a need to carry out new processes to obtain nanoparticles with controlled size and shape, with high performance, high purity, and environmentally friendly conditions and with low levels of cytotoxicity [7,8].

On the other hand, the partial or total substitution of hazardous substances with sustainable processes is known as green chemistry, which arises from the growing demand for more sustainable processes, and is intended to minimize or prevent waste from the reactions involved in the processes of obtaining nanostructured materials while preserving their efficacy [9-11]. This is because, with the obtained characteristics, it is possible to use more biocompatible and renewable chemicals for the development of MNPs with physical, chemical, optical, catalytic and/or electrical properties for their possible application in different areas of science such as medicine [12], biology [13], chemistry [14], and physics [15], among others.

In the search of a nanomaterial for biomedical applications, chitosan (CS), a natural, non-toxic, and biodegradable compound derived from the deacetylation of chitin extracted from the cell walls of some fungi, crustacean exoskeletons and insects, has been studied. [16,17]. Presently, CS is used in the synthesis of gold and silver nanoparticles as a reducing and stabilizing agent at different concentrations of precursor agents, in order to obtain nanomaterials with optical and structural controllable properties, by evaluating with characterization techniques such as UV-visible, infrared spectroscopy by Fourier transform - FTIR, scanning electron microscopy, among others [2].

In addition, the use of gallic acid, present in a variety of plants such as spinach, grapes, green tea, oak bark, red wine and white wine, has been studied as well as chitosan as a reducing agent and stabilizer in the synthesis of gold and silver nanoparticles. In order to compare the catalytic capacity of this material in the process of reducing 4-nitrophenol in the presence of sodium borohydride [18].

Furthermore, the use of ascorbic acid known as vitamin C, present in several fruits and vegetables and used as a dietary supplement, is used as a reducing agent and stabilizer of gold and silver nanoparticles, to evaluate its influence under different conditions of pH for obtaining stable nanostructures and with defined optical properties [19].

On the other hand, the use of sodium alginate, a natural hydrophilic anionic polysaccharide, extracted from brown marine algae [20], which is used in the synthesis of silver nanoparticles as a reducing agent and in turn as a stabilizer of the same, due to this coating provided through the Van der Waals forces. Moreover, its antibacterial effect against

Gram-positive (*Staphylococcus aureus*) and Gram-negative (*Pseudomonas aeruginosa*) microorganisms have been evaluated for possible use in medical devices or wound dressings [21].

Based on the foregoing, this investigation used a non-reducing sugar such as sucrose, a product of the extraction and refining of sugarcane, as a reducing agent of gold and silver nanoparticles, with the aim to evaluate its influence under different volumes of sodium hydroxide to obtain colloidal nanostructures for its possible application in a cellular microenvironment.

2. Methodology

For this investigation, gold trichloride hydrochloride ($\text{HAuCl}_4 \cdot 3\text{H}_2\text{O}$ 99.995 % trace metals basis) was obtained from Aldrich (Sigma-Aldrich), silver nitrate (AgNO_3 pure, pharma grade) was obtained from PanReac (PanReac Applichem). Sodium hydroxide (NaOH pellets) was acquired from Merck (Merck Millipore). Sucrose was obtained at a local market.

2.1. Synthesis of silver nanoparticles (AgNPs)

A 1 mM silver nitrate (AgNO_3) solution was prepared and mixed with 0.2 % sucrose solution. Then, sodium hydroxide (NaOH) 0.1 M was added at volumes of 0 ml, 0.25 ml, 0.5 ml, 0.75 ml and 1.0 ml, the solutions were mixed at 2000 rpm and left to rest at room temperature until a change in color was observed. Finally, the pH reading was performed on each of the samples obtained, using an Ohaus Starter 2100 pH bench meter at standard conditions for temperature and pressure. In addition, blank solutions from the mixture of reducing agent and the different volumes of sodium hydroxide employed in the reaction were prepared. All assays were performed in triplicate.

2.2. Synthesis of gold nanoparticles (AuNPs)

A 0.5mM chloroauric acid (HAuCl_4) was prepared and mixed with 0.1M sucrose solution. Then, 0.1M NaOH was added at volumes of 0 ml, 0.1 ml, 0.15 ml, 0.20 ml and 0.25 ml. Next, the solutions were carried in a water bath at 60 °C by 30 minutes. Once the samples reached standard conditions for temperature and pressure, the pH was read using an Ohaus Starter 2100 pH bench meter. In addition, blank solutions from the mixture of reducing agent and the different volumes of sodium hydroxide employed in the reaction were prepared. All assays were performed in triplicate.

2.3. UV-visible spectrophotometry analysis

Gold and silver nanoparticles were analyzed in a SHIMADZU UV-1800 UV-visible spectrophotometer in a range of 300 - 800 nm; as a control, reducing agent solutions without precursor agent were used. The UV-visible absorbance data obtained for each of the working solutions were subtracted from their respective blank. The result of this subtraction resulted in the resonance associated with the type of metallic particle.

2.4. Fourier transform infrared spectroscopy (FTIR) analysis

The samples were analyzed in an IRAffinity-1S FTIR Spectrophotometer in the range of 500-4000 cm^{-1} to determine the presence of a reducing agent, using attenuated total reflectance (ATR) modulus, with a resolution of 4 cm^{-1} , 128 scans and using Happ-Genzel as apodization function.

2.5. Transmission electron microscopy (TEM) analysis

Analysis by transmission electron microscopy (TEM) was carried out using a Tecnai F20 Super Twin TMP transmission electron microscope operated at 200 kV with the propose to determinate the morphology and particle size.

2.6. Statistical analysis

The synthesis of metal nanoparticles and UV-visible measurements were performed in triplicate in three independent times. In addition, a multivariate statistical analysis was carried out using the Statgraphics Centurion XVI software, which allowed obtaining the population mean, the standard deviation and 95% confidence intervals. For the statistical analysis of the particle size from the TEM images, the ImageJ[®] software was used, obtaining the population mean, the standard deviation and 95% confidence intervals.

3. Results and discussion

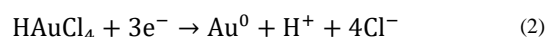
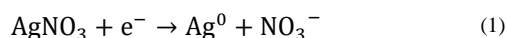
3.1. Non-reducing sugar as a reducing agent in the formation of metal nanoparticles

A reducing sugar is one that through its carbonyl group can donate electrons to molecules that act as oxidants, which accept these electrons. Due to its glycosidic bond, which prevents the opening of the ring of any of its two monosaccharaides, glucose, and fructose, it is not possible to

act as reducing sugar of metal nanoparticles in the absence of hemiacetal groups without a process through which the ring opening of either of the two monosaccharaides [22].

This is achieved by a process of hydrolysis of the sucrose molecule, as shown in Fig. 1, in which the dissolution of this structure into glucose and fructose is generated [23]. At the same time, the glucose and fructose molecules may be present in a cyclic or open form.

The presence of carbonyl groups, especially aldehydes in the case of glucose and ketones for fructose, facilitate the synthesis of metal nanoparticles, where glucose is easier to oxidize than fructose, due to the presence of the hydrogen atom in its main group, becoming a strong reducing agent and responsible for the reduction of gold and silver ions to their metallic states [24]. See eq. (1) and (2).



For a reduction of gold and silver salts to occur, it is necessary that the reducing agent donates electrons, which will be captured by the oxidizing agents. This is known as a reduction-oxidation reaction, in which the structure acting as a reducer, in this case, glucose, passed from an aldehyde group to form a gluconic acid or salt of gluconic acid, related with pH values in the synthesis process (see Fig. 2), an oxidation reaction. While those structures that act as oxidants, gold salts, and silver are reduced and will result in nanometric structures.

In this research, synthesis of nanoparticles of gold and silver by a non-reducing sugar as sucrose was in an alkaline medium, by the addition of different volumes of sodium hydroxide. Upon which, the expected products in the synthesis are salts of gluconic acid because of the oxidation of glucose and the formation of gold and silver nanoparticles. Also, the use of sodium hydroxide allows for the presence of reducing sugars in a short period of time, due to a low dissolution rate of sucrose, requiring the use of acids, alkaline or enzymatic agents, which increase the reaction rate [25,26].

3.2. UV-visible spectroscopy

To demonstrate the formation of gold and silver nanoparticles in the different solutions obtained, these were analyzed by UV-visible spectrophotometry, to identify their respective surface plasmon resonance, which indicates the wavelength in where the greatest electronic movement is presented, translated into an increase in absorbance value [27]. Generally, absorbance peaks have a presence at wavelengths between 400 – 500 nm for silver nanoparticles [27,28] and 500 – 600 nm for gold nanoparticles [29,30]. The absorbance spectra obtained in this investigation are shown below (see Figs. 3 and 4).

From Fig. 3 and 4, it is possible to demonstrate that sucrose in presence of different volumes of sodium hydroxide promotes the formation of gold and silver nanoparticles, from hydrolysis and reduction-oxidation reactions, combined with the presence of the absorbance peaks for each kind of metal between the characteristic wavelength ranges.

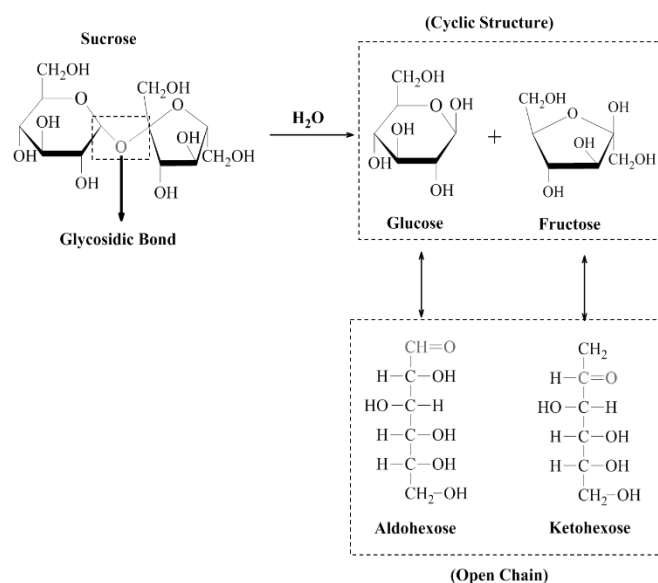


Figure 1. Sucrose hydrolysis into glucose and fructose, in its open and closed form. Source: The authors

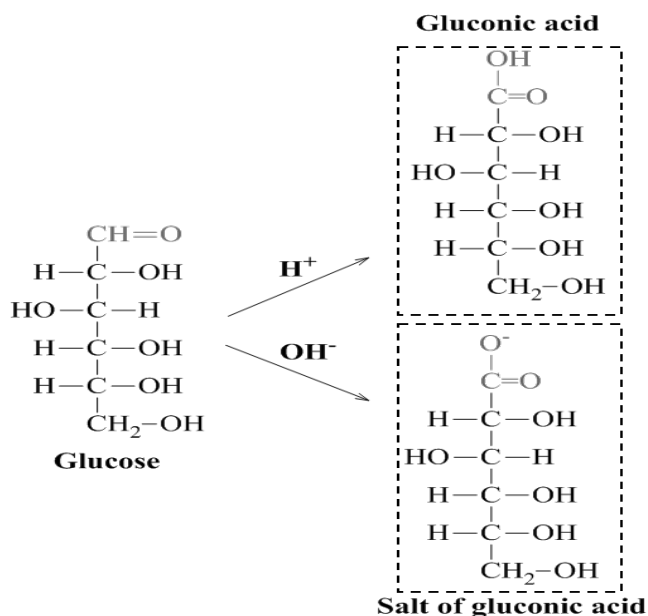


Figure 2. Oxidation of glucose in gluconic acid and gluconic acid salt.
Source: The authors

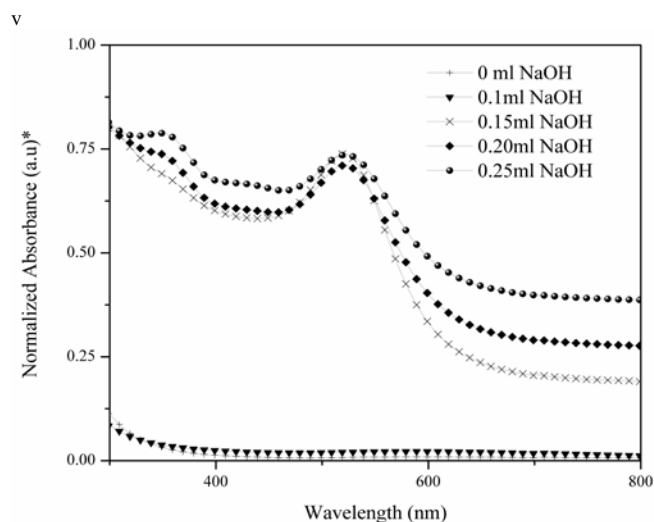


Figure 4. UV-visible spectra of gold nanoparticles at different volumes of sodium hydroxide. As control samples, solutions without reducing agent was used. *The results presented correspond to the population mean of the replicas made. With n = 3
Source: The authors

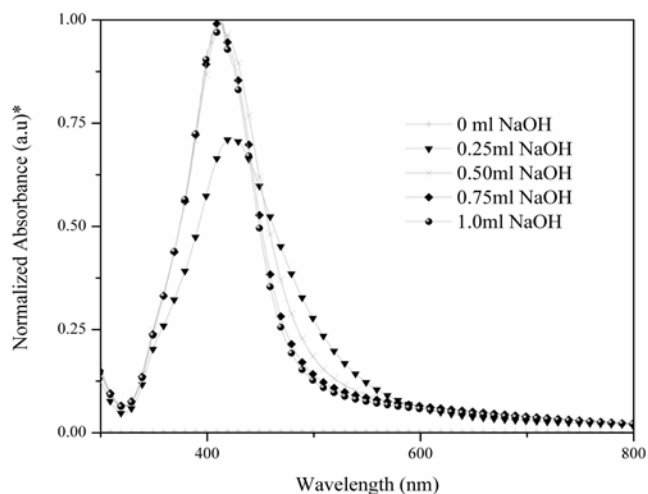
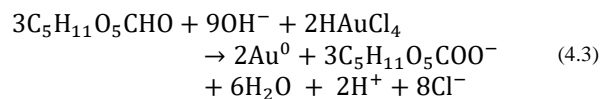
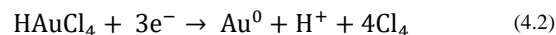
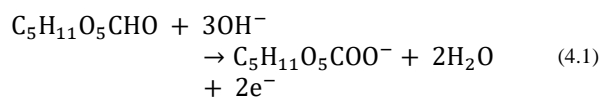
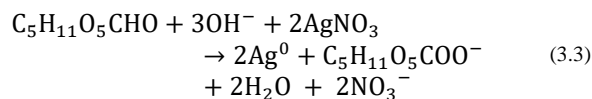
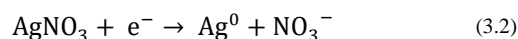
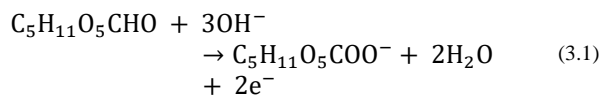


Figure 3. UV-visible spectra of silver nanoparticles at different volumes of sodium hydroxide. As control samples, solutions without reducing agent was used.

*The results presented correspond to the population mean of the replicas made. With n = 3

Source: The authors

In addition, it was possible to show that as of a certain volume of sodium hydroxide is possible to obtain gold and silver nanoparticles. This chemical behavior can be expressed in eq. (3.1) - (3.3) for silver and eq. (4.1) - (4.3) for gold, where is necessary the oxidation of number of glucose moles for the reduction of a particular number of salts precursors moles.



From the above equations and as of the addition of 0.25 ml and 0.15 ml of sodium hydroxide for silver and gold respectively [8], it is possible to obtain nanoparticles (see Figs. 5 and 6). Therefore, the presence of sodium hydroxide besides favoring the formation of silver and gold nanoparticles, at the same time increases the pH values, which significantly decreases the efficiency and formation rate of the nanostructures [31]. This behavior can be observed in the absorption spectra of Fig. 3 and 4, where with volumes of sodium hydroxide of 0.5 ml for silver and 0.15 ml for gold, there is no significant variation in the absorbance values (see Tables 1 and 2).

On the other hand, sodium hydroxide plays an important role in the morphology of nanoparticles; this attribute can be evidenced in TEM images (section 3.4). This was possible to evince it from the behavior of the absorption spectrum of gold, where the presence of a second band (see Fig. 4) at volumes of

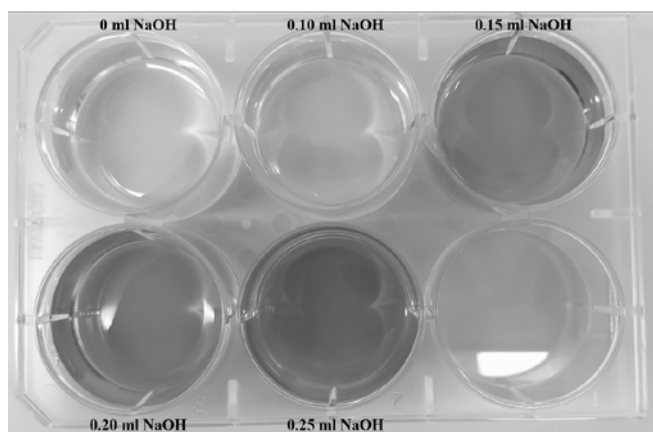


Figure 5. Gold nanoparticles at different volumes of sodium hydroxide. Source: The authors

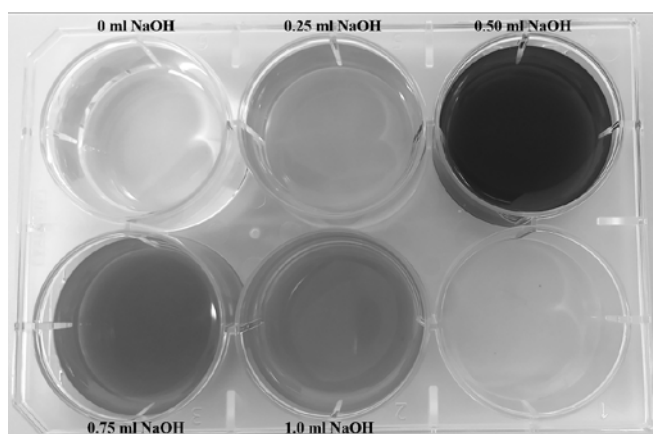


Figure 6. Silver nanoparticles at different volumes of sodium hydroxide. Source: The authors.

Table 1. Wavelengths and absorbance of silver nanoparticles at different volumes of sodium hydroxide.

ml NaOH	pH*	Wavelength SPR (nm)*	Normalized absorbance*	Absorbance (u.a)*
0	6.00 ± 1.76	-	-	-
0.25	9.46 ± 1.43	423 ± 6	0.713±0.268	2.385 ± 0.896
0.50	10.66 ± 1.17	413 ± 5	1.000±0	3.382 ± 0.125
0.75	11.02 ± 0.64	411 ± 5	1.000±0	3.526 ± 0.634
1.00	11.17 ± 0.46	412 ± 2	1.000±0	3.586 ± 0.693

*The data are presented with 95% confidence intervals. With n = 3.

Source: The authors

Table 2. Wavelengths and absorbance of gold nanoparticles at different volumes of sodium hydroxide.

ml NaOH	pH*	Wavelength SPR (nm)*	Normalized absorbance*	Absorbance (u.a)*
0	3.95 ± 0.76	-	-	-
0.1	5.51 ± 0.73	-	-	-
0.15	8.36 ± 1.31	520 ± 2	0.738 ± 0.04	0.195 ± 0.010
0.20	9.49 ± 1.34	522 ± 6	0.710 ± 0.05	0.199 ± 0.010
0.25	10.08 ± 0.51	522 ± 2	0.735 ± 0.04	0.196 ± 0.010

*The data are presented with 95% confidence intervals. With n = 3.

Source: The authors

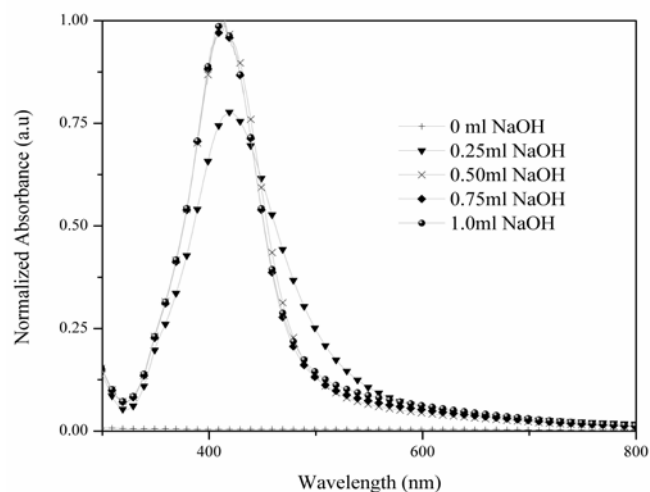


Figure 7. UV-visible spectra of silver nanoparticles at different volumes of sodium hydroxide after four weeks. Source: The authors

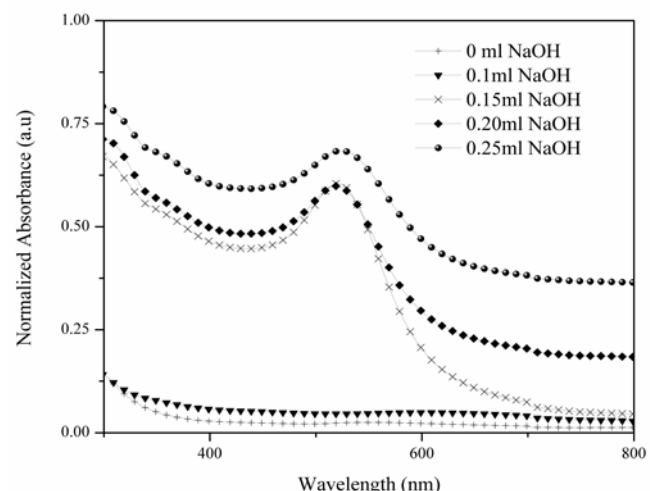


Figure 8. UV-visible spectra of gold nanoparticles at different volumes of sodium hydroxide after four weeks. Source: The authors

sodium hydroxide of 0.20 and 0.25 ml, can present a different morphology to that obtained by using 0.15 ml of sodium hydroxide [32,33]. To confirm the foregoing, the samples were analyzed after four weeks, with the purpose of evidencing changes in the absorption spectra of the gold and silver nanoparticles (see Figs. 7 and 8).

The first findings after four weeks of synthesis was found in the gold and silver absorption spectra, it is the displacement of the superficial plasmon resonance to those initially recorded (see Tables 3 and 4), which can be associated with changes in the colloidal stability of the nanoparticle solution [35]. Second aspect to highlight, is the decrease of the second gold absorption band. Which indicates that high values of sodium hydroxide at the beginning of the reaction, leads to the formation of specific types of morphologies to evidenced two bands in the absorption spectrum, but with the passage of weeks, the reactions become

Table 3. Comparison between wavelengths and absorbance weeks 1 and 4 for silver nanoparticles at different volumes of sodium hydroxide.

ml NaOH	Wavelength week. 1 (nm)	Normalized absorbance week 1	Wavelength week. 4 (nm)	Normalized absorbance week 4
0	-	-	-	-
0.25	424	0.79	420	0.77
0.50	415	1.00	411	1.00
0.75	413	1.00	415	1.00
1.00	413	1.00	415	1.00

Source: The authors

Table 4. Comparison between wavelengths and absorbance weeks 1 and 4 for gold nanoparticles at different volumes of sodium hydroxide.

ml NaOH	Wavelength week. 1 (nm)	Normalized absorbance week 1	Wavelength week. 4 (nm)	Normalized absorbance week 4
0	-	-	-	-
0.1	-	-	-	-
0.15	521	0.72	520	0.60
0.20	520	0.67	519	0.60
0.25	524	0.70	524	0.68

Source: The authors

stabilized in only absorption band could lead to a single morphological structure, this behavior is similar to that reported by Attia et al [32].

3.3. Particle size

From the results obtained in the analysis by UV-visible spectrophotometry, it is possible to determine a particle size approximation from the absorbance and wavelength values obtained for each type of nanoparticle. Using the equations reported by Haiss [35] (5) and Paramelle [34] (6), which are shown below. The results obtained are shown in tables 5 and 6, which are presented as the population mean.

$$dp_{AuNPs} = e^{(3.0 \cdot A_{spr} / A_{450} - 2.2)} \quad (5)$$

$$dp_{AgNPs} = [(\lambda_{max} - 397) / 0.00958]^{0.5} \quad (6)$$

Where A_{spr} and A_{450} correspond to the absorbance obtained at the wavelength of the superficial plasmon resonance and at a wavelength of 450 nm respectively. While λ_{max} is the wavelength of the superficial plasmon resonance. The above equations are valid for a range between [5-80nm] and [20-100nm], respectively.

Additionally, the width at medium height or FWHM was determined, from the UV-vis absorption spectra, to determine experimentally the degree of dispersion of the different synthesized nanoparticles. For this, low FWHM values indicate monodispersity, while high values reflect the polydispersity of particles [36]. The results obtained in this investigation are shown in tables 5 and 6.

Table 5. The experimental particle size and FWHM of silver nanoparticles at different volumes of sodium hydroxide.

ml NaOH	Particle size (nm)	FWHM (nm)*
0	-	-
0.25	52±6	102±19
0.50	41±7	82±12
0.75	38±8	76±8
1.0	40±4	72±7

*The data are presented with 95% confidence intervals. With n = 3. Source: The authors

Table 6. The experimental particle size and FWHM of gold nanoparticles at different volumes of sodium hydroxide.

ml NaOH	Particle size (nm)*	FWHM (nm)*
0	-	-
0.1	-	-
0.15	7 ± 2	126 ± 13
0.20	6 ± 1	142 ± 24
0.25	5 ± 0	161 ± 14

*The data are presented with 95% confidence intervals. With n = 3. Source: The authors

3.4. Transmission electron microscopy analysis

Fig. 9 shows the TEM images obtained for the gold and silver nanoparticles synthesized using sucrose as a reducing agent, in which the morphology, dispersion and average particle size can be evidenced for each type of synthesis performed.

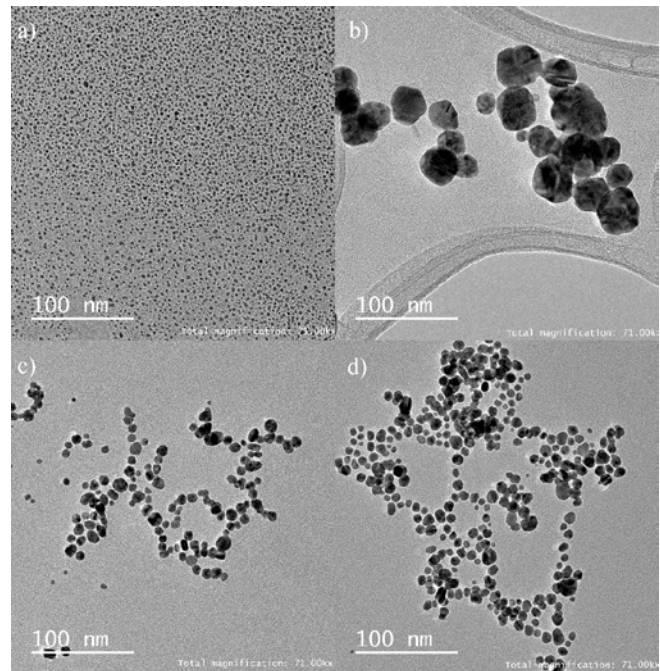


Figure 9. TEM image of metal nanoparticles, a) Silver 0.25 ml NaOH, b) Silver 0.50 ml NaOH, c) gold 0.15 mL of NaOH and d) gold 0.20 ml NaOH. Source: The authors

Table 7.
Particle size analysis obtained by TEM for gold and silver nanoparticles.

Type of nanoparticle	ml NaOH	Particle size range (nm)*	Particle mean size (nm)*
Silver	0.25	2 - 6	3
Silver	0.50	7 - 48	32
Gold	0.15	2 - 15	9
Gold	0.20	3 - 18	9

*The data are presented with 95% confidence intervals. With n = 3.
Source: The authors

Fig. 9 shows the different morphologies obtained from the gold and silver nanoparticles. In the case of silver, when using a volume of sodium hydroxide of 0.25 ml, the initial phase of nanoparticle formation is observed (Figure 9a), while at a volume of 0.50 ml led to formation of structures with a larger diameter and defined morphologies compared at a volume of 0.25 ml (Figure 9b). This behavior is correlated with the results presented in table 7, where average particle diameters of 3 nm are obtained from a volume of sodium hydroxide of 0.25 ml compared to the 32 nm reached at a volume of 0.5 ml.

In relation to gold nanoparticles (Fig. 9c and 9d), greater nanoparticle formation is contemplated as the volume of sodium hydroxide increases, which could be associated

with a greater availability of reducing sugars present in the reaction. The increase in the volume of sodium hydroxide leads to increased polydispersity of particle as shown in Table 6, which relates to the value of width at medium height, where the value changes from 126 nm to a volume of 0.15 ml of sodium hydroxide at 142 nm at a volume of 0.20 ml.

3.5. Fourier-transform infrared spectroscopy analysis

As a complementary technique to UV-vis spectrophotometry transmission electron microscopy, Fourier transform infrared spectroscopy is used to evidence changes in the reducing agent used for the synthesis of metal nanoparticles (see Figs. 10 and 11).

The FTIR spectra of silver nanoparticles exhibit characteristic bands at 3283 and 1637 cm⁻¹, which may be associated with hydroxyl groups either from sucrose, glucose, fructose or gluconic acid [31,37], while the presence of a band at 1738 cm⁻¹ is linked to the stretching of the carbonyl group of the glucose or the gluconic acid. Increasing the volume of sodium hydroxide leads to the appearance and increase of the intensity of the band related to the stretching of the carbonyl group, generated by the breakdown of the glycosidic bond or the oxidation of glucose to gluconic acid.

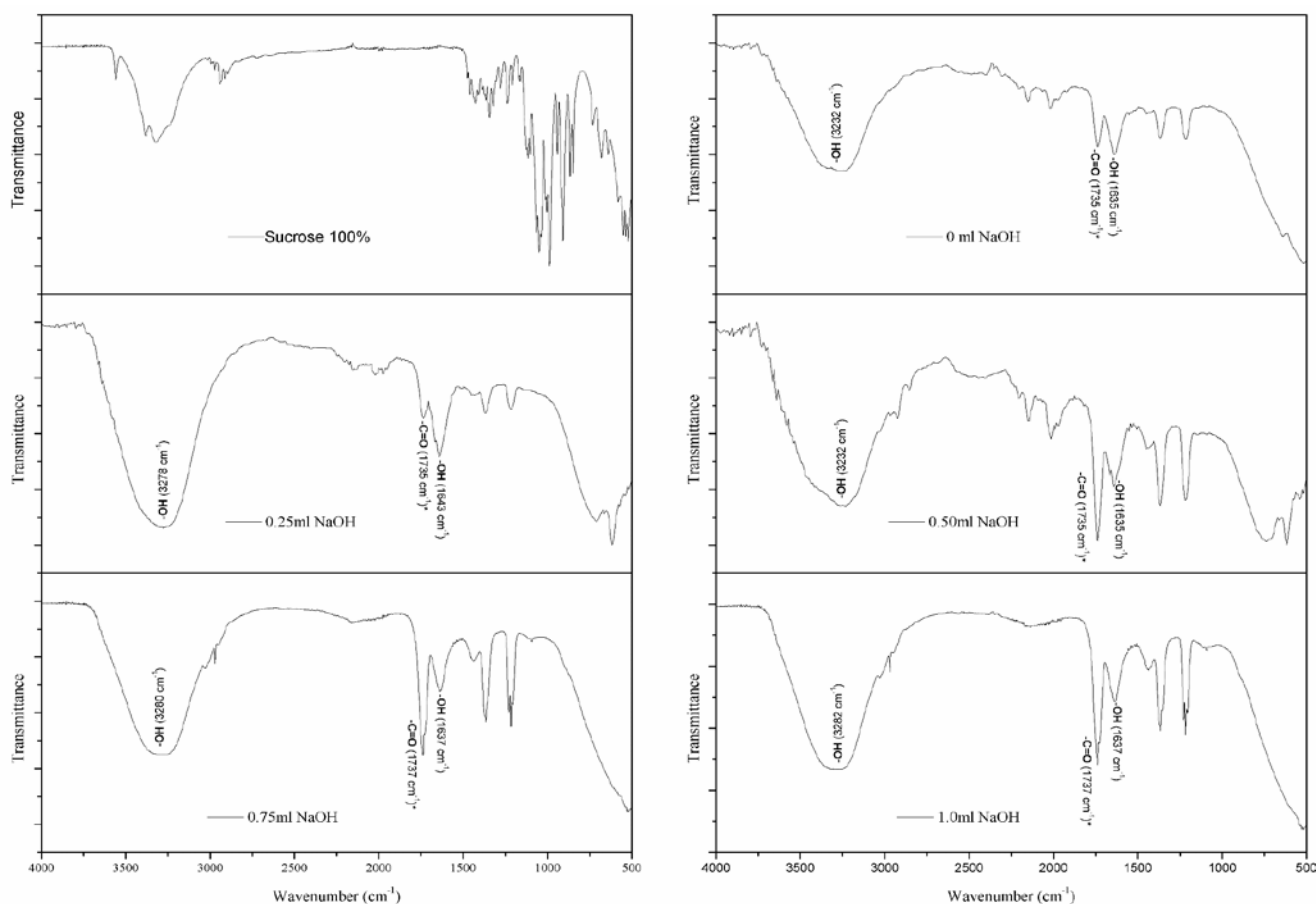


Figure 10. FTIR spectra of silver nanoparticles at different volumes of sodium hydroxide.

*Indicates the change in carbonyl bond intensity (-C=O) as the volume of sodium hydroxide increases in the reaction.

Source: The authors

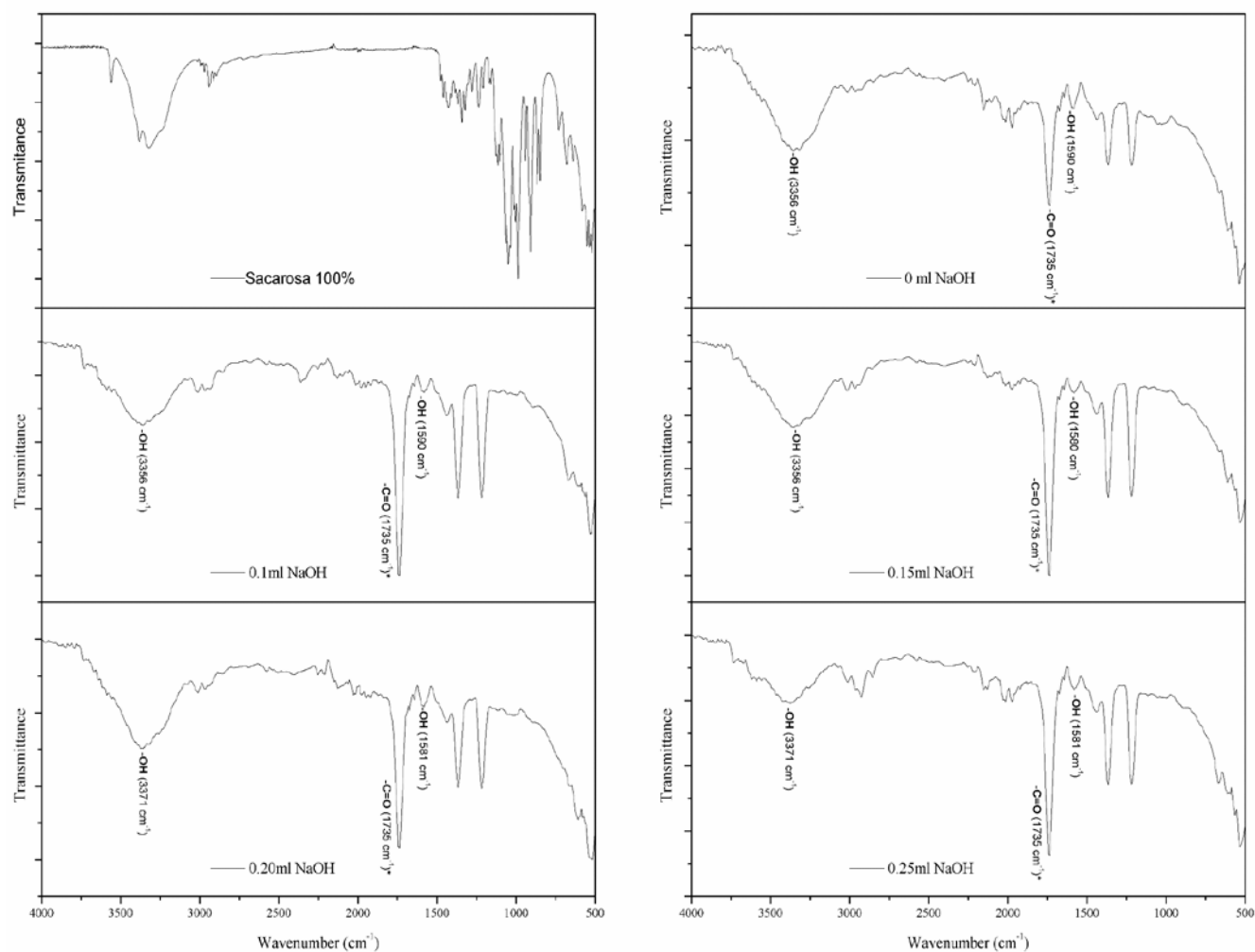


Figure 11. FTIR spectra of gold nanoparticles at different volumes of sodium hydroxide.

*Indicates the change in carbonyl bond intensity (-C=O) as the volume of sodium hydroxide increases in the reaction.

Source: The authors

Moreover, and related with the FTIR spectra of gold nanoparticles (see Fig. 10), it was found that the relative vibration of the hydroxyl group bands are presented at 3371 and 1582 cm^{-1} , while the stretching of the carbonyl group appears at 1736 cm^{-1} . Similarly, to that found in the silver nanoparticles FTIR spectra, the gold nanoparticle FTIR spectra exhibit a higher intensity band associated with stretching of the carbonyl group product of the reduction-oxidation reaction of metal salts and glucose molecule respectively.

4. Conclusion

In this investigation, gold and silver nanoparticles were synthesized by chemical reduction methods using a non-reducing sugar like sucrose, which, through a hydrolysis process leads to structures able of reducing gold and silver salts to its metal states by oxidizing it to salts of carboxylic acid, where these behaviors were evaluated using UV-vis spectrophotometry, TEM and FTIR spectroscopy. Moreover, the effect of sodium hydroxide in the synthesis of MNPs was

evaluated, finding that sodium hydroxides, allows the oxidation of reducing sugars in periods of time less or equal to 30 minutes, but also produces different behaviors related to metal nanoparticles formation rate in relation to the volume of sodium hydroxide used, and particle size obtained, which were determined by using the equations reported by Haiss and Paramelle and the analysis of the images obtained by transmission electron microscopy.

The use of environmentally friendly materials such as sugar offers numerous benefits, ranging from the reduction of environmental risks to the integration of these nanomaterials into biologically relevant systems.

5. Ethical responsibility

Authors declare that no human experiments have been conducted for this investigation. For the investigations related in this document, the approval of the Health Research Ethics Committee of the Universidad Pontificia Bolivariana was obtained by Act dated June 8, 2016.

Author contributions

W. Agudelo worked on the experimental development, design and processing of the spectra and chemical structures of the reducing agent. W. Agudelo, Y. Montoya and J. Bustamante worked on the idea, development and analysis of results, as well as contributing to the writing, review and approval of the content of the manuscript.

Conflict of interest

The authors declare that they have no relation, condition or circumstance that constitutes a potential conflict of interest.

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